Name _____ Date _____ Period _____

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Net Ionic Equations (Single)

Honors Chemistry

Use the activity series to predict the products and write the balanced molecular equation with states. Split every (aq) compound into its ions to write the total ionic equation, and keep elements, solids, liquids, and gases together. Cross out the spectator ions, write the net ionic equation, and list the spectators. Assume iron forms Fe^{2+}.

1. $Mg_{(s)} + CuCl_{2(aq)} \rightarrow$

Molecular:

Total ionic:

Net ionic:

Spectator ions:

2. $Zn_{(s)} + HCl_{(aq)} \rightarrow$

Molecular:

Total ionic:

Net ionic:

Spectator ions:

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3. $Fe_{(s)} + CuSO_{4(aq)} \rightarrow$

Molecular:

Total ionic:

Net ionic:

Spectator ions:

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4. $Cl_{2(g)} + KBr_{(aq)} \rightarrow$

Molecular:

Total ionic:

Net ionic:

Spectator ions:

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